

## PERIODIC TRENDS PRACTICE

Use the periodic trends organizer to answer the following questions.

### Atomic Radius practice

1. For each pair of elements, predict which atom is larger.

A. Mg, Sr.

D. Ge, Br

B. Sr, Sn.

E. Cr, W

C. Ge, Sn

### Ionization Energy practice

2. For each of the following pairs, predict which atom has the higher ionization energy.

A. Mg, Na

D. Cl, I

B. S, O.

E. Na, Al

C. Ca, Ba

F. Se, Br

### Electronegativity practice

3. For each of the following pairs, predict which atom has the higher electronegativity.

A. Mg, Na

D. Ca, Ba

B. Na, Al

E. S, O

C. Cl, I

F. Se, Br

**Ionic Radius** (The portion of the organizer that says “cation” and “anion”.)

4. For each of the following pairs, predict which atom or ion is larger.

A. Mg, Mg<sup>2+</sup>

D. Cl<sup>-</sup>, I<sup>-</sup>

B. S, S<sup>2-</sup>

E. Na<sup>+</sup>, Al<sup>3+</sup>

C. Ca<sup>2+</sup>, Ba<sup>2+</sup>

F. Mg<sup>2+</sup>, S<sup>2-</sup>