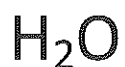


Try to draw some of these structures. Follow the steps until you get used to it.



N		Lewis Dot Structure	<ol style="list-style-type: none">1. Fill out NASL2. Choose the atom that goes in the middle.3. Use the shared electrons to bond the other atoms to the middle.4. Fill in lone pairs where atoms need them for an octet.
A			
S			
L			



N		Lewis Dot Structure	<ol style="list-style-type: none">1. Fill out NASL2. Choose the atom that goes in the middle.3. Use the shared electrons to bond the other atoms to the middle.4. Fill in lone pairs where atoms need them for an octet.
A			
S			
L			



N		Lewis Dot Structure	<ol style="list-style-type: none">1. Fill out NASL2. Choose the atom that goes in the middle.3. Use the shared electrons to bond the other atoms to the middle.4. Fill in lone pairs where atoms need them for an octet.
A			
S			
L			

Cl_2 (2 chlorine atoms, not carbon and 2 iodine)

N		Lewis Dot Structure	<ol style="list-style-type: none">1. Fill out NASL2. Choose the atom that goes in the middle.3. Use the shared electrons to bond the other atoms to the middle.4. Fill in lone pairs where atoms need them for an octet.
A			
S			
L			

CH_2O

N		Lewis Dot Structure	<ol style="list-style-type: none">1. Fill out NASL2. Choose the atom that goes in the middle.3. Use the shared electrons to bond the other atoms to the middle.4. Fill in lone pairs where atoms need them for an octet.
A			
S			
L			

HCN

N		Lewis Dot Structure	<ol style="list-style-type: none">1. Fill out NASL2. Choose the atom that goes in the middle.3. Use the shared electrons to bond the other atoms to the middle.4. Fill in lone pairs where atoms need them for an octet.
A			
S			
L			