

Naming acids (compounds that start with H):

- Binary Acids – Two elements; one is always hydrogen. For acids without oxygen, the name is written as “hydro[anion]ic acid”.
 - H_2S = “hydrosulfuric acid”
 - HBr = “hydrobromic acid”
 - HCN = “hydrocyanic acid”
- For acids that contain oxygen (called “oxyacids”): The name of the acid is “[anion name][suffix] acid.”
 - The suffix depends on the name of the anion:
 - If the anion ends in “-ate”, the suffix is “-ic”.
 - HNO_3 = nitric acid
 - H_2SO_4 = sulfuric acid
 - H_3PO_4 = phosphoric acid
 - If the anion ends in “-ite”, the suffix is “ous”.
 - HNO_2 = nitrous acid
 - H_2SO_3 = sulfurous acid
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The Table below will help you with prefixes and suffixes

Type	Binary	Oxyacid
Ion ends in	-ide	-ate
Changed ending to	-ic	-ic
Add to front	Hydro	
Add to End	Acid	Acid
Example	HI	$\text{H}_2(\text{SO}_4)$
Name	Hydroiodic acid	Sulfuric acid

Naming Bases is Easy!

Write the name of the metal then write the word hydroxide. Strong bases result from combining group 1 or group 2 metal with hydroxide.

Ammonia, NH_3 is a base that doesn't have a metal or hydroxide. It's an exception to memorize.